

SCANNING ELECTROCHEMICAL IMPEDANCE MICROSCOPY FOR THE EVALUATION OF BIOLOGICAL SURFACES

Inga Morkvėnaitė-Vilkončienė^{1,2}, Antanas Zinovičius^{1,3}, Justė Rožėnė¹, Arūnas Ramanavičius^{3,4}

¹Department of Mechatronics, Robotics, and Digital Manufacturing, Vilnius Gediminas Technical University, J. Basanavičiaus 28, Vilnius, Lithuania.

²Laboratory of Electrochemical Energy Conversion, State Research Institute Centre for Physical Sciences and Technology, Saulėtekio 3, Vilnius, Lithuania

³Department of Physical Chemistry, Faculty of Chemistry and Geosciences, Vilnius University, Naugarduko 24, Vilnius, Lithuania

⁴Laboratory of Nanotechnology, State Research Institute Centre for Physical Sciences and Technology, Saulėtekio 3, Vilnius, Lithuania

*inga.morkvenaite-vilkonciene@vgtu.lt.

Scanning electrochemical microscopy (SECM) is a powerful tool for the evaluation of reaction kinetics, occurred at the surface of interest [1]. Using SECM, it is possible to determine the surface's electrochemical activity by scanning it with ultramicroelectrode (UME) in a standard electrochemical cell. By using bipotentiostat, the surface of interest can be connected to the electrochemical cell. Potentiostat, used in SECM allows to apply available electrochemical techniques, such as amperometry, cyclic voltammetry, potentiometry, electrochemical impedance spectroscopy (EIS) and so on. If EIS is applied with SECM, the electrochemical system is called 'scanning electrochemical impedance microscopy' (SEIM). SEIM could be used for the registration of electrochemical impedance spectroscopy locally at the surface of interest [2,3]. This technique is beneficial in corrosion researches, but also can be applied for the evaluation of biological surfaces (Fig. 1). The application of SEIM for the assessment of biological surfaces is the main goal of our research.

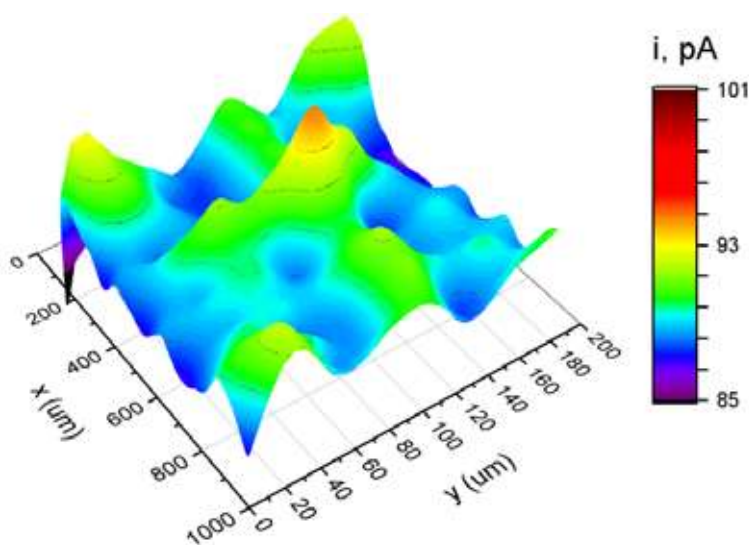


Fig. 1. Mapping of electrochemical activity of horseradish peroxidase, immobilized at 100 μm spots. Measurement was performed in buffer solution, pH 6.7 with 1 mM H_2O_2 .

References

- [1] I. Morkvenaite-Vilkonciene, P. Genys, A. Ramanaviciene, A. Ramanavicius, *Col. Sur. B.* **2015**, 126, 598.
- [2] A. Valiūnienė, J. Petronienė, I. Morkvenaite-Vilkonciene, G. Popkirov, A. Ramanaviciene, A. Ramanavicius, *Phys. Chem. Chem. Phys.* **2019**, 21, 9831.
- [3] I. Morkvenaite-Vilkonciene, A. Valiūnienė, J. Petronienė, A. Ramanavicius, *Electrochem. Commun.* **2017**, 83, 110.

Acknowledgements

This research was funded by the European Social Fund according to the activity "Development of Competencies of Scientists, other Researchers and Students through Practical Research Activities" measure Project No .09.3.3.-LMT-K-712-16-0211.